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7(c)	$2\text{MnO}_4^- + 6\text{H}^+ + 5\text{H}_2\text{O}_2 \rightarrow 2\text{Mn}^{2+} + 8\text{H}_2\text{O} + 5\text{O}_2$ <p>Moles $\text{MnO}_4^- = (24.35/1000) \times 0.0187 = \underline{4.55 \times 10^{-4}}$</p> <p>Moles $\text{H}_2\text{O}_2 = (4.55 \times 10^{-4}) \times \underline{5/2} = 1.138 \times 10^{-3}$</p> <p>Moles H_2O_2 in 5 cm^3 original $= (1.138 \times 10^{-3}) \times \underline{10} = 0.01138$</p> <p>Original $[\text{H}_2\text{O}_2] = 0.01138 \times \underline{(1000/5)} = 2.28 \text{ mol dm}^{-3}$ (allow 2.25-2.30)</p>	1 1 1 1 1	if no equation and uses given ratio can score M2, M3, M4 & M5 Note value must be quoted to at least 3 sig. figs. M2 is for 4.55×10^{-4} M3 is for $\times 5/2$ (or $7/3$) Mark consequential on molar ratio from candidate's equation M4 is for $\times 10$ M5 is for consequentially correct answer from (answer to mark 4) $\times (1000/5)$ Note an answer of between 2.25 and 2.30 is worth 4 marks) If candidate uses given ratio $3/7$ max 4 marks: M1: Moles of $\text{MnO}_4^- = \underline{4.55 \times 10^{-4}}$ M2: Moles $\text{H}_2\text{O}_2 = (4.55 \times 10^{-4}) \times \underline{7/3} = 1.0617 \times 10^{-3}$ M3: Moles H_2O_2 in 5 cm^3 original $= (1.0617 \times 10^{-3}) \times 10 = 0.01062$ M4: Original $[\text{H}_2\text{O}_2] = 0.01062 \times (1000/5) = 2.12 \text{ mol dm}^{-3}$ (allow 2.10 to 2.15)
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